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## **Balancing Equations and Simple Stoichiometry-KEY**

For example, acetic acid will be mixed with sodium hydroxide

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solution. Upon mixing, the weak acid and hydroxide ion react to produce acetate ions and water. As long as hydroxide ion is the limiting reagent, and some acetic acid remains, the solution contains both species of the conjugate pair (acetic acid and acetate ions), and a buffer exists.

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### Worksheet 1 Answer

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9) What is the limiting reagent for problem 8? oxygen

10) How much of the excess reagent is left over after the reaction from problem 8 is finished? 21.5 grams of  $C_6H_{10}$  will be left over.

11) If 35 grams of carbon dioxide are actually formed from the reaction in problem 8, what is the percent yield of this reaction? 80.1%

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