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Write answers in the boxes below. 23. What is the mass of 0.5 mole of ammonia,  $\text{NH}_3$ ? 24. What is the mass of 2 moles of ?  $\text{Ca}(\text{OH})_2 + 3242$  25. What is the mass of 4.2 moles of  $\text{H}_3\text{PO}_4$ ? 26. What is the mass Of 0.75 mole of butane,  $\text{C}_4\text{H}_{10}$  911,6 47 50 27. What is

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the mass of 2 moles of gypsum,  
 $\text{CaC}_2\text{O}_4 \cdot 2\text{H}_2\text{O} \times 2(35) 4$

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## Grant Proposal Assessment

Silberstein, 2003 Element Subscript  
Oxidation state Sum of oxidation states  
K Cr +2 TOTAL -14 9. 10. 11. [a]  
potassium is a group one metal; its  
oxidation state is always +1 [b] oxygen  
usually has an oxidation state of -2 [C]  
the sum of oxidation states of each  
element is the product of the subscript  
and the oxidation state

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