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Limiting Reactants In your textbook, read about why reactions stop and how to determine the limiting reactant. Study the diagram showing a chemical reaction and the chemical equation that represents the reaction.

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CHAPTER Section 11.1 continued In your textbook, read about mole ratios.

Answer the questions about the following chemical reaction. sodium + iron(III) oxide  $\rightarrow$  sodium oxide + iron

$6\text{Na}(s) + \text{Fe}_2\text{O}_3(s) \rightarrow 3\text{Na}_2\text{O}(s) + 2\text{Fe}(s)$  15. What is a mole ratio? ... Study Guide Chemistry: Matter and Change Chapter 11 ...

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TEACHER GUIDE AND ANSWERS Study Guide - Chapter 11 - Stoichiometry

Section 11.1 What is stoichiometry? 1. true 2. true 3. false 4. true 5. true 6. 2, 2, 64.10 7. 3, 3, 96.00 8. 2, 2, 88.02 9. 4, 4, 72.08 10. methanol and oxygen gas 11. carbon dioxide and water 12. 160.10 g 13. 160.10 g 14. They are equal. 15. A mole ratio is a ratio between ...

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12 . ... Chapter 12 Stoichiometry . In the reaction represented by the equation  $2\text{Na} + 2\text{H}_2\text{O} \rightarrow 2\text{NaOH} + \text{H}_2$ , how many grams of hydrogen are produced if 120. g of Na and 80.0 g of  $\text{H}_2\text{O}$  are available? . For the reaction represented by the equation

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Block Scheduling Lesson Plans

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### **CHAPTER CHAPTER 12 Study ...**

Chapter 11 study guide True/False

Indicate whether the statement is true

or false. \_\_\_\_ 1. The actual yield is

always lower than the theoretical yield.

Matching Match each item with the

correct statement below. a.

stoichiometry b. mole ratio c.

stoichiometric equivalent \_\_\_\_ 2.

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Section 11.1 What is stoichiometry? ...  
12. 13. 14. 72 Methanol Oxygen gas



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Carbon dioxide Water What are the  
reactants? C) What are the products?  
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