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Chapter 11 Chemical Reactions Answers

Chemistry chapter 11 Chemical reactions answer key. coefficient. a whole number that appears before a formula in an equation. spectator ion. a particle not directly involved in a chemical reaction. combustion reaction. a reaction in which oxygen reacts with another substance, often producing light or heat. reactant.

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Chemistry Chapter 11 Chemical Reactions Worksheet Answers

oxidation-reduction reactions A chemical reaction that exhibits a change in the oxidation states of one or more elements in the reactants that has the general form oxidant + reductant → reduced oxidant + oxidized reductant. The general forms of these five kinds of reactions are summarized in Table 11.6.1, along with examples of each.

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Section 11.3 - Reactions in Aqueous Solution. A complete ionic equation is an equation that shows dissolved ionic compounds as dissociated free ions. A spectator ion is an ion that is not directly involved in a reaction. Compounds that are aqueous will break into ions, and compounds that are solid will remain bonded.

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Chapter 11: Chemical Reactions Study Guide chemical equation A representation of a chemical reaction with reactants on the left, products on the right, and an arrow separating the two skeleton

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11.1 Describing Chemical Reactions Worksheet Answers

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Answer Key Chapter 11: Chemical Reactions Chemical Equations Questions 1. Early Chinese alchemists discovered an early form of gunpowder. What was the composition of this substance? 2. What later developments were made to the gunpowder recipe that improved its pyrotechnic properties? 3.

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Chapter 11: Chemical Reactions

Write correct formulas for the products in these double replacement reactions. $\text{EH}_3\text{PO}_4 + 2\text{K}_2\text{CO}_3 + \text{BaCl}_2 \rightarrow \text{---} + \text{A} + \text{IN}_3 + \text{Co}_5\text{AgNO}_3 + \text{KCl} + \text{CJ} + \text{kH}_2\text{SO}_4 \rightarrow + \text{O} + \text{CoQ} + 7\text{BC}_2\text{H}_3\text{O}_2 + 3\text{kI} + \text{O} + \text{Poq} + 3\text{CaS} + 50 + 361 + 10\text{AgC}_2\text{H}_3\text{O}_2 + \text{K}_2\text{CrO}_4$

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