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In a strong acid - strong base titration, neutralization produces water and an aqueous solution of a salt, whose cation and anion come from the base and acid, respectively. Neither ion is acidic or basic, so the pH at the equivalence point

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is that of neutral water; i.e., 7.00.

Worksheet 22 Titrations Key - Illinois

Worksheet 23 - Strong Acid/Strong Base
Titrations A. Initial pH This is always
determined based solely on the initial
concentration of the acid or base being
titrated. Every mole of acid or base will
produce one mole of H^+ or OH^- .

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Exception: Bases formed with Group II cations will form two moles of OH⁻ for every mole of base, e.g. Ca(OH)₂.

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Student Support Services Program 1) It takes 83 mL of a 0.45 M NaOH solution to neutralize 235 mL of an HCl solution. What is the concentration of the HCl solution? 2) You are titrating an acid into a base to determine the concentration of the base. The

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**Acid Base Titrations Chem
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Worksheet 22 - Weak Acid/Strong Base

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Titrations A. Initial pH This is determined by the initial concentration of the weak acid and the dissociation constant, K_a .
 $[H^+][A^-] = K_a [HA]$
Set up an ICE table for the dissociation of the weak acid in water. The “change” line

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