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Water And Aqueous Systems Chapter 15 Chemistry - ProProfs Quiz

aqueous solution: a

solution in which the

solvent is water:

solvent: the dissolving

medium in a solution:

surfactant: wetting

agent that interferes

with hydrogen bonding

in water: strong

electrolyte: a

substance that

completely dissociates

into its ions in solution:

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water of hydration: the
water loosely held in a
crystal structure:
Brownian motion

Chemistry Chapter 15: Water & Aqueous Systems Flashcards ...

Chapter 15 water and
aqueous systems. 1.
Chapter 15 "Water and
Aqueous Systems" Pre-
AP Chemistry Charles
Page High School
Stephen L. Cotton. 2.
Section 15.1 Water and

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it's Properties

OBJECTIVES: -Explain the high surface tension and low vapor pressure of water in terms of the structure of the water molecule and hydrogen bonding.

Chapter 15, Water and Aqueous Systems - Guided Practice ...

Chapter 15 - Water and Aqueous Systems -
15.3 Heterogeneous Aqueous Systems -

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15.3 Lesson Check -
Page 507: 21 Answer
Colloids cannot be separated through filtering because their particles are smaller than those of suspensions and they do not settle out over time.

SECTION 15.1

WATER AND ITS PROPERTIES (pages 445-449)

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Chemistry Chapter 15:

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describes a substance that removes sufficient water from the air to form a solution; the solution formed has a lower vapor pressure than that of the water

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in the air. suspension.
a mixture from which
some of the particles
settle out slowly upon
standing.

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the essential lessons

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associated with water
and aqueous systems.

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Chapter 15 (Water and Aqueous Systems) Test - Chapter 15 ...

Chemistry chapter 15
water and Aqueous
systems Vocab. An
inward force that tends
to minimize the surface
area of a l... Any
substance that
interferes with
hydrogen bonding
between wa... The

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dissolving medium in a solution Surface tension An inward force that tends to minimize the surface area of a l... Surfactant Any substance...

Chapter 15 - Water and Aqueous Systems - 15.2

Homogeneous ...

Water And Aqueous Systems Chapter 15 Chemistry. In a full glass of water, the water surface seems to

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bulge over the rim of the glass. Water forms nearly spherical drops at the end of an eyedropper. An insect called a water strider is able to "walk" on water.

Chapter 15 Water and Aqueous Systems Worksheet Answers ...

CHAPTER 15, Water and Aqueous Systems(continued) 6. Circle the letter next to

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each sentence that describes a result of the surface tension of water. a. In a full glass of water, the water surface seems to bulge over the rim of the glass. b. Water beads up into small, nearly spherical drops on a paper towel.

Chapter 15 - Water and Aqueous Systems - Standardized Test ...

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Chemistry: Chapter 15: Water and Aqueous Systems ...

Chapter 15 - Water and
Aqueous Systems -
15.2 Homogeneous
Aqueous Systems -
15.2 Lesson Check -
Page 501: 17 Answer
CH₄ doesn't dissolve in
water because it is a
molecular compound
with no net dipole KCl
is soluble because of
its ionic bonds.

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chapter 15
Answers
chemistry aqueous
systems ... - Quizlet

Chapter 15, Water and
Aqueous Systems -
Guided Practice

Problem? GUIDED

PRACTICE PROBLEM 6.

Calculate. c. Determine
the mass of water in
the hydrate. mass of
 $5\text{H}_2\text{O} = 5 \times [(2 \times \underline{\quad}) +$
 $\underline{\quad}] = 5 \times \underline{\quad} = \underline{\quad} \text{ g.}$

d. Determine the mass
of the hydrate.

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05 CTR ch15 7/12/04

8:14 AM Page 387

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Chemistry Chapter

15: Water and

Aqueous ...

The high surface tension of water is due to the:

- small size of water molecules.
- low mass of water molecules.
- hydrogen bonding between water molecules.
- covalent bonds in water molecules.

_____ 12.

Salts and other

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compounds that remove moisture from air are said to be: a. efflorescent. c. colloidal. b. surfactant. d. hygroscopic. 10 9 ...

chemistry chapter 15 water aqueous systems Flashcards and ...

The Water Molecule: a Review Water is a simple tri-atomic molecule, H_2O Each O-H bond is highly polar, because of the high

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electronegativity of the oxygen (N, O, F, and Cl have high values) bond angle of water = 105° due to the bent shape, the O-H bond polarities do not cancel. This means: water is a polar molecule.

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Chapter 15: Water and Aqueous Systems. It takes more energy to break the ion

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attractions between water molecules. If the hydrogen bonding in water was weak it would be a gas at the usual temperatures.

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Chapter 15 - Water and Aqueous Systems - 15.3

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Heterogeneous ...

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“Water and Aqueous Systems”

Chapter 15 (Water and Aqueous Systems) Test Study Guide The bonds between the hydrogen and oxygen atoms in a

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water molecule are polar covalent bonds. The covalent bonds are polar because the oxygen atom has a greater electronegativity than the hydrogen atoms. The bonds between adjacent water molecules are called hydrogen bonds.

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systems worksheet
answers , source:trajan
scimed.com. Informal
together with feedback
sessions help do away
with minor splinters
that may hamper the
practice of achieving
the vision. Adhere to
the instructions about
what to edit. The
estimating worksheet
is designed to direct
you.

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